

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Subatomic Particles for Atoms, Ions and Isotopes-Practice

Use your periodic table to complete this worksheet.

	Atom/Ion/Isotope name	Standard Atomic Notation	Atomic Mass	Atomic Number	Number of Protons	Number of Electrons	Charge	Number of Neutrons
1	Strontium (isotope)		90					
2				4			+2	
3		${}^{79}_{34}\text{Se}^{2-}$						
4	Aluminum (Ion)							14
5			3		2			
6			75				0	
7					3			8
8				11		10		
9		${}^{201}_{80}\text{Hg}$						
10			53		26			
11	Nitrogen (ion)						-3	
12		${}^{59}_{28}\text{Ni}^{3+}$						
13	Chlorine (isotope)		36					
14	Xenon (atom)							
15					30			29

**\*NOTE:** when you are finished, you should have 3 atoms, 6 ions and 6 isotopes.